* made of ions
* many soluble in water but not in nonpolar liquid
* made of molecules
* poor electrical conductors
* Many soluble in nonpolar liquids but not in water
* Elements from opposite ends of the periodic table will form this kind of bond
* LARGE differences in electronegativity
* Elements are close together in electronegativity values
* CO2
* H2O
* NaCl
* C6H12O6
* attraction between oppositely charged ions in a chemical compound
* These kinds of bonds occur mainly between a metallic and a non metallic atom.
* bonding between two non metallic atoms which is characterized by the sharing of pairs of electrons between atoms
* Chemical bonds
* Prefixes used when naming
* Carbon tetrachloride
* Binary compound name ends in –ide
* Lithium acetate



* 
* Nonpolar
* Polar
* Involve electrons
* Involve nonmetals
* Involve metals
* Donating electrons to become a positive ion
* Cations
* Anions
* Representative particles are formula units (f.u.)

 

* Lewis dot diagrams show these with straight lines
* Lewis dot diagrams show these with an arrow moving an electron
* chemical substance consisting of two or more different chemically bonded chemical elements, with a fixed ratio determining the composition
* The flame test lab used compounds of this type
* one atom gains electrons to form an anion
* satisfy the octet rule
* electrons are shared equally
* H2
* electrons are shared unequally
* dipole interactions